

**PH102, 2012W, Lecture Notes: March 2, Fri, Class 19**

**Nuclear Physics: Nuclear Structure, Properties, and Models**

**Structure of Nuclei**

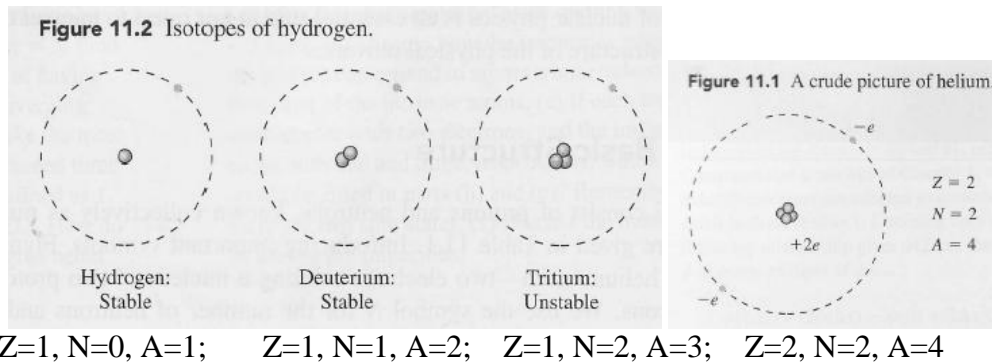
Nuclei consist of nucleons, i.e. protons and neutrons.

- $Z$ = number of protons
- $N$ = number of neutrons
- $A$ =mass number=total number of nucleons =  $Z + N$

	charge	Mass in kg	Mass in u
Proton	+e	$1.6726217 \times 10^{-27}$	1.007276
Neutron	0	$1.6749273 \times 10^{-27}$	1.008665
Electron	-e	$9.109 \times 10^{-31}$	0.0005486

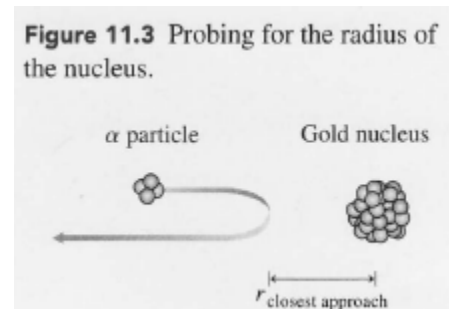
Where 1u (atomic mass) = 1/12th of the mass of C-12

- Isotopes= Nuclei with the same number of protons but different numbers of neutrons. Isotopes that are not present in nature are unstable. Some isotopes that are abundant in nature are unstable.
- Hydrogen isotopes
  - Hydrogen’s nucleus has one proton and is stable
  - Deuterium’s nucleus has one proton and one neutron and is stable, 0.015% of the hydrogen atoms found in nature are Deuterium.
  - Tritium’s nucleus has one proton and two neutrons and is unstable.



**Size of Nucleus**

- Rutherford’s experiment with  $\alpha$  particles (=He’s nuclei after electrons are removed) fired toward the thin Gold film.
  - The nuclei are roughly spherical.
  - Most of alpha particles do not scatter
  - The head-on approach would enable alpha particles to approach the gold nuclei closest. The experiment showed that  $r = A^{1/3}R_0$  where



$$R_0 = 1.2 \times 10^{-15} \text{ m.}$$

- The nuclear dimension is about  $10^{-15}$  meter (femto-meter)

- Nuclear volume

$$V = \frac{4}{3} \pi r^3 = \frac{4}{3} \pi (A^{1/3} R_0)^3 = \frac{4}{3} \pi A R_0^3 \sim A \text{ (proportional to mass number)}$$

- Nuclear density is about the same for all nuclei since volume is proportional to mass (number).

$$\text{density} = \frac{\text{mass}}{\text{volume}} = \frac{A \times \text{mass of nucleon}}{A \times \frac{4}{3} \pi A R_0^3} \cong 10^{17} \text{ kg/m}^3$$

- Nucleons in a nucleus is closely packed and each nucleon's volume in a nucleus is

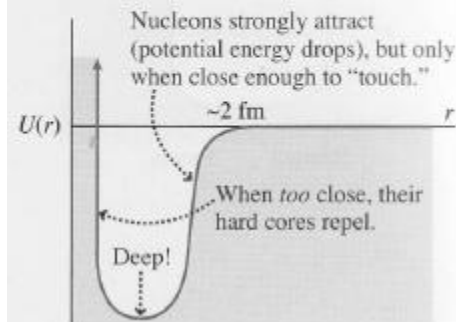
$$\frac{V}{A} = \frac{4}{3} \pi R_0^3$$

Thus, the effective radius of a nucleon is  $R_0$ .

### Strong Force:

- Nucleons inside a nucleus are held together by strong force
- Strong force is attractive
- Strong force is short-ranged (about 2 fm) and strong
- Strong force is nearly identical between n-n, n-p, and p-p.
- The potential energy for strong force is shown in Figure 11.4.

**Figure 11.4** The basic elements of the internucleon (strong force) potential energy—a strong, short-range attraction with a repulsive hard core.



### Comparison of forces

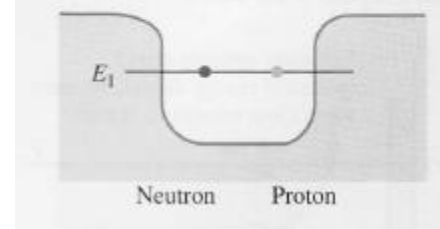
Force	Applying to	Relative strength	Range
Strong	Between nucleons	1	$\sim 1 \text{ fm}$
Electromagnetic	Between charges	$\sim 10^{-2}$	$\propto 1/r^2$ (long range)
Weak	Related to radioactive decay	$\sim 10^{-6}$	$\sim 10^{-3} \text{ fm}$
Gravitational	Between masses	$\sim 10^{-39}$	$\propto 1/r^2$ (long range)

**Nuclear binding:** not all nucleon arrangements are stable.

### Two-nucleon nuclei:

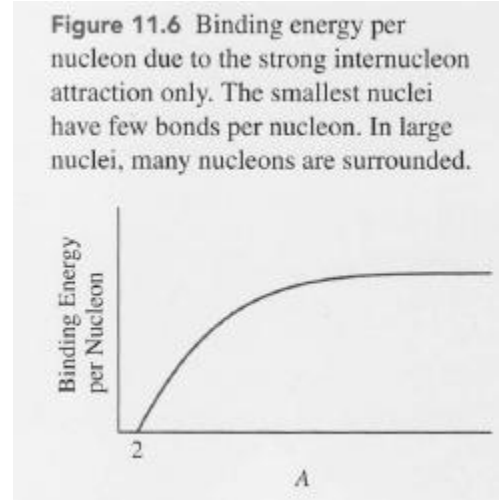
- Protons and neutrons are fermions and are subject to the Exclusion Principle.
- Parallel spin arrangements are more stable.
- Proton-neutron can have a parallel spin arrangement in the ground spatial state while proton-proton and neutron-neutron cannot do this.
- Proton-Neutron thus is stable and has one bound state. And Deuterium's nucleus has a total spin of 1.

**Figure 11.5** The deuteron's neutron and proton bound in a well resulting from their attractive potential energy.

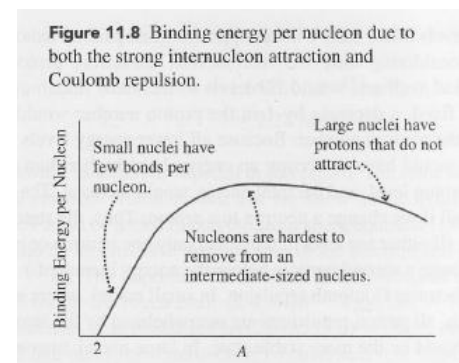
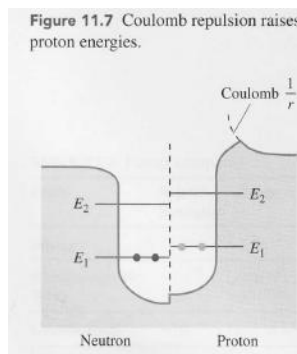


Arbitrary number of nucleon nuclei:

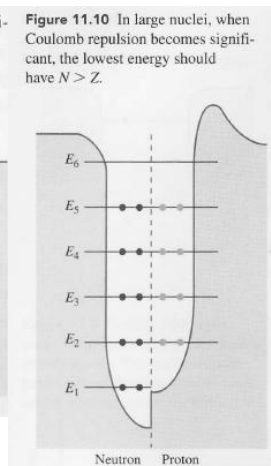
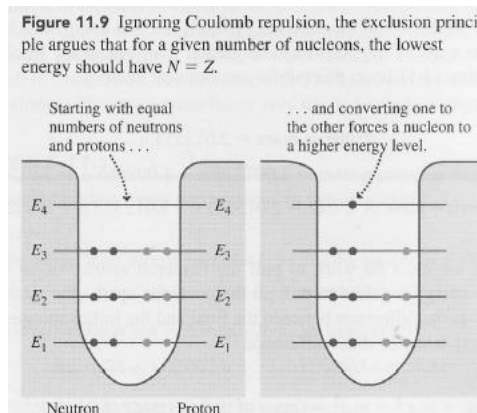
- Strong force
  - As more and more nucleons are present in nuclei, the number of bonds among nucleons increases, e.g. (2 nucleons  $\rightarrow$  1 bond, 3 nucleons  $\rightarrow$  3 bonds, 4 nucleons  $\rightarrow$  6 bonds, etc.). This indicates that the binding energy of a nucleus will increase as the number of nucleons increases.
  - However, the strong force is short ranged and attracting only the nearby nucleons, making each surrounding nucleon has the same number of bonds. Nucleons at the surface are not completely surrounded. The ratio of surface nucleons over total number of nucleons diminishes as  $A$  grows  $\propto 1/r$ . Figure 11.6 shows this trend.



- Coulomb repulsion
  - Proton-proton will repel if they are not in contact with one another. See Figure 11.7 where proton's energy levels are slightly elevated due to repulsive force.
  - Coulomb force becomes more and more act as a destabilizing force as nuclei get larger.



- The Exclusion Principle
  - Protons and neutrons obey the exclusion principle independently.
  - For smaller nuclei, the lowest energy state can be obtained with equal numbers of protons and neutrons. See Figure 11.9
  - For larger nuclei, the lower energy state can be obtained when  $N > Z$  due to the increase in repulsive force between protons. See Figure 11-10.



### Liquid Drop Model for Nucleus:

Binding energy = volume term + surface term + Coulomb term + Asymmetry term

$$= C_1 A - C_2 A^{2/3} - C_3 \frac{Z(Z-1)}{A^{1/3}} - C_4 \frac{(N-Z)^2}{A}$$

where

$$C_1 = 15.8$$

$$C_2 = 17.8$$

$$C_3 = 0.71$$

$$C_4 = 23.7$$

The binding energy formula above is called the semiempirical binding energy formula because the formula is based on theoretical argument with empirical evidence.

Example: Fe-56

$$\text{Binding energy} = 15.8 \times 56 - 17.8 \times 56^{2/3} - 0.71 \frac{26 \times 25}{56^{1/3}} - 23.7 \frac{(30-26)^2}{56} = 496.9 \text{ MeV}$$

$$\text{Binding energy/nucleon} = 496.9 \text{ MeV}/56 = 8.87 \text{ MeV}$$

Figure 11.16 shows the actual data with the semiempirical binding energy formula plot. Fit is good except very light nuclei such as He's nucleus.

### Binding Energy

$$\text{Binding energy} = (\text{mass of individual nucleons} - \text{mass of nucleus})c^2$$

Example: Deuteron's binding energy

Deuteron = 1 proton + 1 neutron

$$\begin{aligned} \text{Binding Energy} &= (\text{proton mass} + \text{neutron mass} - \text{Deuteron mass})c^2 \\ &= (1.007276 \text{ u} + 1.008665 \text{ u} - 2.013553 \text{ u})c^2 = 0.002388 \text{ u}c^2 \\ &= 0.002388 \times 1.661 \times 10^{-27} \text{ kg} \times (3 \times 10^8 \text{ m/sec})^2 = 2.22 \text{ MeV} \end{aligned}$$

Compare the binding energy of electron in the hydrogen atom (=13.6 eV) and the binding energy of nucleons in the deuterium  $\sim 10^{-5}$

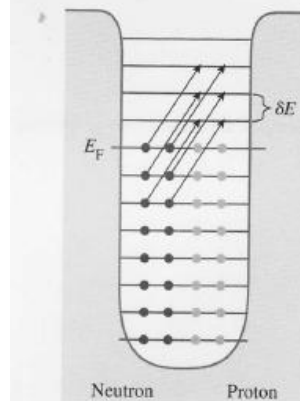
$$\text{Binding Energy} = (Zm_H + Nm_n - M_{A_X}^X)c^2$$

Where  $m_H$  = atomic mass of hydrogen

$m_n$  = neutron mass

$M_{A_X}^X$  = atomic mass of the nucleus

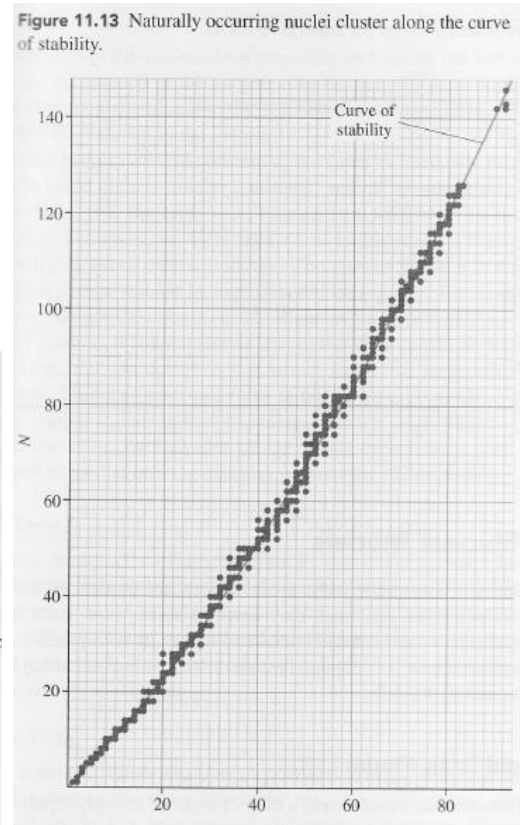
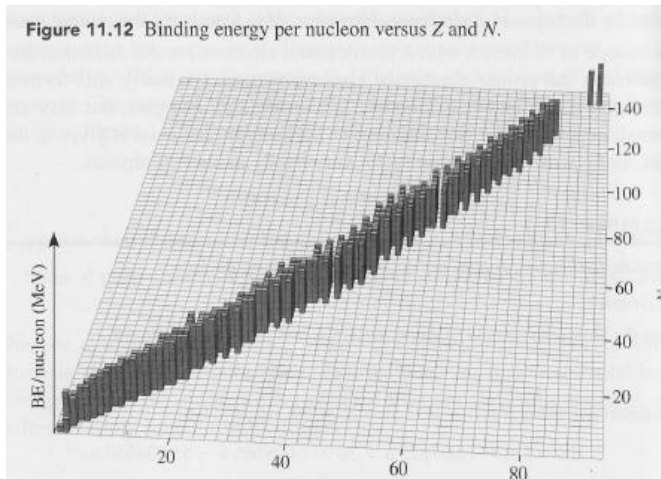
Figure 11.15 If  $j$  neutrons become protons, the energy increases by  $\frac{1}{2}j^2\delta E$ .



For example:  ${}^{56}_{26}\text{Fe} = 55.934939 u$

Binding energy =  $(26 \times 1.007825 u + 30 \times 1.008665 u - 55.934939 u) c^2 = 0.528461 uc^2 = 492.3 \text{ MeV}$  (consider  $c^2 = 931.5 \text{ MeV}/u$ ).

The curve of stability: Binding energy/nucleon for naturally occurring isotopes  
Figure 11.12 and Figure 11.13



Binding energy per nucleon over  $A$ : Figure 11.14

